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Chapter 11. Chemical Reactions Ch 11 Guided Reading Section 11.1 Describing Chemical Reactions p. 346-354 . By the end of this section you should be able to: write skeleton and formula equations; write and balance chemical equations. Notes: Ch 11.1 Ch 11.1 Handouts:

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Chapter 11 Chemical Reactions 119 GUIDED PRACTICE PROBLEM 2 (page 324) 2. Sulfur burns in oxygen to form sulfur dioxide. Write a skeleton equation for this chemical reaction.

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Include appropriate symbols from table 11.1. Analyze Step 1. Identify the relevant concepts. Write the formula for each reactant and each product. Include the common

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321–329)

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Chapter 11: Chemical Reactions Study Guide ?chemical equation A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two skeleton

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Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride. $\text{CuS (s)} + \text{O}_2\text{(g)} \rightarrow \text{Cu (s)} + \text{SO}_2\text{(g)}$ D

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Chapter 11: Chemical Reactions Study Guide. A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two. Small whole number ratios that are placed in front of the formulas in an equation in order to balance it. Chapter 11: Chemical Reactions Study Guide | StudyHippo.com

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Section 11.1 Describing Chemical Reactions 323 with ChemASAP Chemical Equations Word equations adequately describe chemical reactions, but they are cumbersome. It's easier to use the formulas for the reactants and products to write chemical equations. A chemical equation is a representation of a chemical reaction; the formulas of the reactants (on

11.1 Describing Chemical Reactions 11

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Stoichiometry Problems. When we carry out a reaction in either an industrial setting or a laboratory, it is easier to work with masses of substances than with the numbers of molecules or moles. We will first present the method used in most other books for converting from the mass of any reactant or product to the mass of any other reactant or product using a balanced chemical equation outlined ...

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