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chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

## SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process.

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Introduce the term sto- ichiometry in your own words.  
Stress that stoichiometry allows students to calculate  
the amounts of chemical sub- stances involved in

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chemical reactions using information obtained from balanced chemical equations.

## 12.1 The Arithmetic of Equations 12

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Chapter 3: Stoichiometry – Guided Reading Section

3.1 – 3.2 1. True or False? Most hydrogen atoms have a mass of 1.008 amu. Justify your answer. If true, explain why it is true. If false, what mass do most hydrogen atoms have? False, 1.008 amu is actually



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hydrogen's average mass, NO atom of hydrogen actually has the mass of 1.008 amu. 2.

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