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Chapter 16 – FIN 3403 – Practice Problem Solutions 2 EBIT Indifference Point Since we are going to need all of the values eventually, here is a table with Net Income in each State (with and without debt):

	State 1	State 2	State 3	State 4	Debt No Debt	Debt No Debt	Debt No Debt
Debt No Debt EBIT	\$150,000	\$150,000	\$120,000	\$120,000	\$80,000	\$80,000	\$40,000
Interest					-\$14,000	\$0	-\$14,000
EBT	\$136,000	\$150,000	\$106,000	\$120,000	\$66,000	\$80,000	\$26,000
Taxes							

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Chapter 16 Section 16.2 (continued) Sample Problem 16.3 Answers 10. $335 \text{ mL} \times 1\text{L}/1000 \text{ mL} \times 0.425 \text{ mol}/1\text{L} = 1.42 \times 10^{-1} \text{ mol}$ 11. $0.2 \text{ mol}/\text{L} \times 1\text{L}/1000 \text{ mL} \times 250 \text{ mL} = 5.0 \times 10^{-1} \text{ mol}$ CaCl₂; $5.6 \times 10^1 \text{ g}$ CaCl₂ Practice Problems Plus How many grams of solute are in 2.40 L of 0.650 M HClO₂? (107 g) Math Handbook For a math refresher and practice, direct students to conversion

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Solutions to the Extra Problems for Chapter 16 1. The equation is $CS_2 + 4H_2 \rightarrow 2H_2S + CH_4$. To solve this problem, we just have to think backwards from Equation (16.1). The exponents are the coefficients in the chemical equation, and the products are in the numerator of the equation for K, while the reactants are in the denominator. 2.

~~Solutions to the Extra Problems for Chapter 16~~

Practice-Problem-6660020 book December 22, 2010 8:37 Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of $(NH_4)_2SO_4(aq)$ is given by $x_{(NH_4)_2SO_4} = \frac{n_{(NH_4)_2SO_4}}{n_{(NH_4)_2SO_4} + n_{H_2O}}$ Because $(NH_4)_2SO_4(aq)$ is a strong electrolyte, it dissociates completely into $NH_4^+(aq)$ and $SO_4^{2-}(aq)$ ions. Assume a one kilogram solution.

~~CHAPTER 16. Colligative Properties of Solutions~~

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CHAPTER 16- Practice Exercise A. Kweyete . Acid- Base Equilibria and Solubility Equilibria. Problem #1. What would be the effect on pH by adding the following to a solution of NH_3 : (a) NH_4Cl Decrease pH (b) NaCl NO EFFECT (c) NaOH Increase pH (d) HCl Decrease pH

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