

Chapter 17 Study Guide Acids Bases

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~~Chapter 17~~

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Acids are named by changing the ending of the alkane to -oic acid. i.e. a 4 carbon acid goes from butane to butanoic acid Acid salts are named by changing the acid ending to -oate i.e the 4 carbon acid goes from butanoic acid to sodium butanoate Esters are named by changing the ending of the alkane to -oate and by using alkyl name i.e. the 4 carbon ester below is methyl butanoate(methyl for OCH₃ portion) O O Anhydrides are dehydrated esters that replace -oic acid with word anhydride Acyl ...

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Carboxyl, amine, hydrogen and variable side group Describe two forces that help shape an amino acid chain? Peptide bond holds chains of amino acids together by linking amine and carboxyl groups; hydrogen bonds help fold proteins into shapes and hold them together. Shape depends on amino acids, polarity and R groups.

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•The most common strong acids are HCl, HBr, HI, HNO₃, HClO₃, HClO₄, and H₂SO₄. •Strong acids are strong electrolytes. •All strong acids ionize completely in solution. •Example: Nitric acid ionizes completely in aqueous solution. HNO₃(aq) + H₂O(l) → H⁺(aq) + NO₃⁻(aq) •Since H⁺ and H₃O⁺ are used interchangeably, we write HNO

~~AP Chemistry CHAPTER 16 STUDY GUIDE Acid Base Equilibrium~~

There are 20 amino acids, but only 4 nucleotide bases in DNA. That means it takes at least 3 nucleotides to code for one amino acid, hence a codon

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(three amino acids). ? The code is read as an RNA code, 5' to 3'.

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