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ox: $[\text{Ni}(s) \rightarrow \text{Ni}^{2+} + 2e^{-}] \times 5$; red: $[5e^{-} + 8\text{H}^{+} + \text{MnO}_4^{-} \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}]$

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4 → Mn²⁺ + + 4H₂O] × 2.

balanced: 5Ni(s) +

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(aq) → 2Ag + (aq) + Cu(s) E

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+) = - 0.80V + 0.34V = -
0.46V. Not spontaneous. ox:
[Ni(s) → Ni²⁺ + 2e⁻] × 5;
red: [5e⁻ + 8H⁺ + MnO₄⁻
→ Mn²⁺ + 4H₂O] × 2.
balanced: 5Ni(s) + 16H⁺
(aq) + 2MnO₄⁻(aq) → 5Ni²⁺
(aq) + 2Mn²⁺(aq) + 8H₂O

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20.

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Chapter 18 - Nuclear

Chemistry 289 Key Ideas

Answers 14. Because protons and neutrons reside in the nucleus of atoms, they are called nucleons. 16. There

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are two forces among the particles within the nucleus. The first, called the electrostatic force, is the force between electrically charged particles. The second force,

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a) the energy given off when reactants collide b) *an energy barrier between reactants and products the heat released in a reaction* c) generally very high for a reaction that takes place

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Chapter 18 Amines and

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Amides. 18.1 Multiple Choice Questions. 1) The compound $\text{CH}_3 - \text{CH}_2 - \text{NH} - \text{CH}_3$ is classified as a. A) primary amine. B) secondary amine. C) tertiary amine. D) quaternary amine. E) hydrated amine. Answer: B.

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Objective: 18.1. Global

Outcomes: GO2. 2) The

compound $\text{CH}_3 - \text{CH}_2 - \text{NH}_2$ is classified as a. A) primary amine. B) secondary amine.

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