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Structure of

atom [?][?][?][?][?] [?][?]

[?][?][?][?][?][?] part 1

The Chemicals of

Life Unit 1.4 -

Composition of

Mixtures [?][?][?][?][?][?]

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parts) 66 terms.

eline. Leçon 17.

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2] 42 terms.

eline. Chapter

4: Sensation and

...

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Introduction to
the Chemistry of
Life. Figure 2.1
Foods such as
bread, fruit,
and cheese are
rich sources of
biological
macromolecules.

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The elements
carbon,
hydrogen,
nitrogen,

oxygen, sulfur,
and phosphorus
are the key
building blocks
of the chemicals
found in living
things. They
form the
carbohydrates,
nucleic acids,

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proteins, and
lipids (all of
which will be
defined later in
this chapter)
that are the
fundamental
molecular
components of
all organisms.

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the Chemistry of

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Chapter 2 The

Life Chemistry Of

Chapter 2 The
Chemistry of
Life What are

the basic
building blocks
of all matter?
Atoms Describe
the structure of
an atom protons,
electrons, and
neutrons. The
nucleus (center)
of the atom

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contains the
protons
(positively
charged) and the
neutrons (no
charge) .

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Question: What
Answers

are the basic
chemical
principles that
affect living
things?

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Chapter 2 The

of Life. Atom.

Nucleus.

Electron.

Element. An atom

is the smallest
constituent unit
of ordinary

matter th... The

nucleus is the

small, dense

region

consisting of

protons a... The

electron is a

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subatomic
particle, symbol
e⁻ or β^- , with
a....

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A B; atom: the
basic unit of
matter: nucleus:
the center of

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the atom: Chemistry Of

electron: a Life Worksheet

negatively

charged Answers

particle:

element: a pure

substance that

consists

entirely of one

type of atom

Quia - Chapter

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of Life

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chemistry.

Chemistry is the study of matter, its properties, how and why substances combine or separate to form other substances, and

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how substances
interact with
energy. 2.

matter. The term
matter refers to
anything that
occupies space
and has mass—in
other words, the
“stuff” that the
universe is made
of. 3. three
different levels
of matter - The

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Chapter ...

A compound is a substance formed by the chemical combination of two or more elements in definite proportions.

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(this means that

H_2O is water,
but H_3O is

not). Compounds

have different

physical and

chemical

properties from

the elements

they are made of

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2, The Chemistry

of Life

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Discovery of Electron-Discharge Tube Experiment
In 1879, William Crooks studied the conduction of electricity through gases at low pressure. He

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performed the
experiment in a
discharge tube
which is a

cylindrical hard
glass tube about
60 cm in length.
It is sealed at
both the ends
and ...

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show 1 [...]
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Life What do you
see when you
look at this
picture? Is it

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just a mass of tangled ribbons? Look closely.

It's actually a complex pattern of three-dimensional shapes. It represents the structure of a common chemical found inside living cells.

The chemical is

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a protein called
kinase.
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2 Part 2. 68

terms.

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terms.

meganbutler6211.

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Chapters 1, 2, 3,

& 4) 48 terms.

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Preparation, and
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Properties of
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Answers

2. What relationship exists between the mass number of an element and isotopes of that element? 3.

Explain the difference between ionic and covalent bonds. 4.

Compare and contrast

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adhesion and
cohesion. Below
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synthesis.

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Subatomic
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particles that
Answers
make up the at...

The atomic
number. A
chemical
reaction in
which two
molecules are
bonded
together... A
chemical process
that splits a

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molecule by
adding water.
Protons,
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number of
protons.

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looks at atoms,
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molecules. Good
review of
chemistry we see
in microbiology.

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2-4 Completion

On the lines

provided,

complete the

following

sentences. 1.

Chemical

reactions that

energy often

occur

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spontaneously.

2. During a
chemical
reaction,

chemical bonds
are 3.

Biological
catalysts, or
enzymes, act by
lowering the
required for a
reaction. 4.

Biochemistry 11

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Section 2-3

Carbon Compounds
(Pages 44-48)

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