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Chapter 3 Review Atoms Section
CHAPTER 3 REVIEW Atoms: The
Building Blocks of Matter SECTION 3
SHORT ANSWER Answer the
following questions in the space
provided. 1. Explain the difference
between the mass number and the
atomic number of a nuclide. Mass
number is the total number of protons
and neutrons in the nucleus of an
isotope.

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3 Atoms: The Building Blocks of Matter

1. All matter is composed of atoms. 2. Atoms of a given element contain all the same properties, and vice versa. 3. Atoms cannot be created, destroyed, or subdivided. 4. Atoms of different elements combine in simple whole numbered ratios, to form chemical compounds. 5. In chemical reactions, atoms are combined, separated. or rearranged.

chemistry chapter 3 review atoms

Flashcards | Quizlet

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CHAPTER 3 REVIEW . Atoms: The Building Blocks of Matter. SHORT ANSWER . Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the . cathode 2. The smallest unit of an element that can exist either alone or in molecules containing.the

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter
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_____ . 2.

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1. Elements are made of particles that are very small known as atoms. 2. All of the atoms of the same element are duplicates of each other/identical. 3. The atoms of one element are different from those of any other type of element. 4. Atoms of one element can combine with atoms of different elements to make compounds.

Chemistry- Chapter 3, Review Questions: Section 2 ...

3. State two principles from Dalton's atomic theory that have been revised as new information has become available. 4. The formation of water

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according to the equation. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. shows that 2 molecules (made of 4 atoms) of hydrogen and 1 molecule (made of 2 atoms) of oxygen produce 2 molecules of water.

CHAPTER 3 REVIEW -

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Chapter Four [Arrangement of Electrons in Atoms] Chapter Five [The Periodic Law] Chapter Six [Chemical Bonding] ... Section 2: Chapter review 3 and 6. Section 3: Chapter review 7 thru 16. Work practice problems 17, 18, 19, 21,22 and 23 (pp 89-90). Homework Answers . Problem Solving Packet.

Chapter Three [The Building Blocks of Matter]

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Acces PDF Chapter 3 Review Atoms Section 1 Periodic Table Atoms: The Building Blocks of Matter SECTION . 3. SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the . mass number. and the . atomic number. of a nuclide. 2. Why is it necessary to use the average atomic mass of all isotopes, rather than the

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Chapter 4 Arrangement Of Electrons In Atoms Section 3

CHAPTER 4 REVIEW. Arrangement of Electrons in Atoms. SECTION 4.3. SHORT ANSWER Answer the following questions in the space provided. 1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. 2. Explain the conditions under which the following orbital notation for helium is ...

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Section Quiz, continued Class Date

extra of atoms with how many 6.

Bohr's model correctly explains the electron(s)? a. one b. two c. three d. four

or more 7. A form of energy that

exhibits wave behavior as it travels

through space is a. microwave

radiation. b. ultraviolet radiation. c.

infrared radiation. d. all of the above

8.

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CHAPTER 5 Electrons in Atoms

Resource Manager Objectives Section

Section 5.1 What You'll Learn Using

the Photo Electrons in Atoms Have

students review the following concepts

Analysis Answers will vary.

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