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Chapter 3: Pure Substances – Thermodynamics
The change in internal energy can be found from the first law of thermodynamics: U = Q-W = (3.5 x 10 J) - (2.1 x 10 J) = 0.9 x 10 J = 90 kJ.) A gas in a cylinder is kept at a constant pressure of 3.5 x 10 Pa while 300 kJ of heat are added to it, causing the gas to expand from 0.9 m³ to 1.5 m³ .

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3-5 3-23 Problem 3-22 is reconsidered. The missing properties of water are to be determined using EES, and the solution is to be repeated for refrigerant-134a, refrigerant-22, and ammonia.

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Chapter 3.THERMODYNAMICS. -Thermodynamicsis the study of the relationship between the energy transformationin the system and other physical quantities such as temperature, pressureand volume (P, V, T). -A thermodynamic equation of stateis a mathematical relationship of the thermodynamic or state variables, such as pressure, volume and temperature.

Chapter 3: THERMODYNAMICS
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Chapter 3-3 Heat transfer is energy in transition due to a temperature difference. The three modes of heat transfer are conduction, convection, and radiation. Conduction through Plane Walls Conduction heat transfer is a progressive exchange of energy between the molecules of a substance. Fourier's law of heat conduction is QAk(dT cond t dx = – here Q!

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