

## Chapter 6 Thermochemistry Review Answers

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2 CHAPTER 6 THERMOCHEMISTRY 7. In calorimetry, heat flow is determined into or out of the surroundings. Because  $\Delta E_{univ} = 0$  by the first law of thermodynamics,  $\Delta E_{sys} = -\Delta E_{surr}$ ; what happens to the surroundings is the exact opposite of what happens to the system.

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Chapter 6 Thermochemistry Review Answers Thermochemistry Exam1 and Problem Solutions 1. Which ones of the following reactions are endothermic in other words  $\Delta H$  is positive? I.  $\text{H}_2\text{O}(\text{l}) + 10,5\text{kcal} \rightarrow \text{H}_2\text{O}(\text{g})$   $\Delta H_1$  II.  $2\text{NH}_3 + 22\text{kcal}$  Thermochemistry Exam1 and Problem Solutions | Online ...

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### ~~Chapter 6 Thermochemistry Mrs. Duffey FHN~~

AP Chemistry Practice Test, Ch. 6: Thermochemistry Name \_\_\_\_\_ MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be \_\_\_\_\_ and has a \_\_\_\_\_  $\Delta H$  at constant pressure. A)endothermic, positive B)endothermic, negative C ...

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PDF CHAPTER 6 THERMOCHEMISTRY - Oregon State University 162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write:  $101.3 \text{ J } 0.18 \text{ L atm } 1 \text{ L atm} = - \square \times = \square \text{ w } -18 \text{ J}$  6.17 An expansion implies an increase in volume, therefore  $w$  must be  $-325 \text{ J}$  (see the defining equation for pressure-volume work.)

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6. Answer the following questions with regards to the Heating Curve. 1) What is happening to the average potential energy of the molecules in the sample during section 3? Potential Energy remains constant. It only increases or decreases when temperature is remains constant on the plateaus. ... Thermochemistry Review Worksheet ...

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162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write:  
 $101.3 \text{ J} \cdot 0.18 \text{ L atm} \cdot 1 \text{ L atm}^{-1} = -18 \text{ J}$   
6.17 An expansion implies an increase in volume, therefore  $w$  must be  $-325 \text{ J}$  (see the defining equation for pressure-volume work.) If the system absorbs heat,  $q$  must be  $+127 \text{ J}$ . The change in energy (internal energy) is:

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Equilibrium. The balanced equation for this reaction is:  $2 \text{C}_8\text{H}_{18}(\text{l}) + 25 \text{O}_2(\text{g})$ .  
Next Answer Chapter 6 - Thermochemistry - Review Questions - Page 284: 2  
Previous Answer Chapter 5 - Gases - Marathon Problems - Page 244: 165. Acevedo was born in 1975, so subtract that year from the current year to find her age.

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Chemistry Chapter 6 Notes Thermochemistry Daniel Seo 10/10/13 Period 5 Chapter 6 Thermochemistry Outline 6.1 The Nature of Energy 1. Energy ? capacity to do work (or to produce heat) i. Work ? force acting over a distance (moving an object) ii. Heat ? form of energy. ? chemicals may store potential energy in their bonds that can be released as heat energy 2.

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