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OBJECTIVES: – Determine the number of valence electrons in an atom of a representative element. 3. Section 7.1 - Ions OBJECTIVES: – Explain how the octet rule applies to atoms of metallic and nonmetallic elements. 4.

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206 Chapter 7 • Ionic Compounds and Metals Section 7.1.1 Figure 7.1 As carbon dioxide dis-solves in ocean water,

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Carbonate ions are produced. Coral polyps capture these carbonate ions, producing crystals of calcium carbonate, which they secrete as an exo-skeleton. Over time, the coral reef forms. A coral reef is a complex habitat that

Chapter 7: Ionic Compounds and Metals

Chapter 7 Ionic and Metallic Bonding

Multiple Choice Identify the choice that

best completes the statement or answers

the question. _____ 1. What are the valence

electrons of an atom? a. The electrons that

an atom gains when it becomes an anion

b. The electrons that remain closest to the

nucleus c. The electrons in the highest

occupied energy level d. The electrons that

an atom loses when it ...

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Lesson 7.1 Reading and Study Workbook

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Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged. Lesson Summary. Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation. For a representative element, the group number equals the ...

BONDING AND INTERACTIONS

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bonds Lewis structures and the octet rule
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and electronegativity Formal charge
Chapter 8 Energetics of bond formation
and dissociation Bond formation: $A + B$
 AB If the reaction is energetically
favorable, AB is lower in energy than $A +$
 B . Heat is released .

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