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Consists of a single

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atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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single atom with a
positive or negative

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charge resulting
from the loss or
gain of electrons,
respectively.

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Naming and
Writing Formulas
for Acids and
Bases. An acid is a

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compound that produces H^+ ions when it dissolves in water. The formula for an acid normally starts with an H. When naming acids, you should first determine whether or not the acid contains oxygen.

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Acids. base. Cation.

Chemical
Ionic compounds.

Names And
Compounds that

Formulas
contain one or

more hydrogen

atoms and

Answers
produce.... An ionic

compound that

produces

hydroxide ions

when dissolved....

Any atom or group

of atoms that has a

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positive charge.

Compounds
composed of
cations and anions.

Formulas

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SECTION 9.3
NAMING AND

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Chapter 9
WRITING
FORMULAS FOR
MOLECULAR
COMPOUNDS
(pages 268–270)

This section
explains the rules
for naming and
writing formulas for
binary molecular
compounds.

Naming Binary
Molecular
Compounds (pages

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268-269) 1. Circle
the letter of the
type(s) of elements
that form binary
molecular
compounds.

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~~CHEMICAL NAMES~~
~~AND FORMULAS 9~~
~~SECTION 9.2~~
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Chapter 3 IONIC COMPOUNDS

1. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following binary ionic compounds. sodium sulfide g. CuCl: h. snc.l. a.

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Mn02 c. d. SrBr2

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Formulas

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“Chemical Names
and Formulas” ... in
parenthesis after
the name of the

metal (Table 9.2,
p.255) Predicting
Ionic Charges

Some of the post-
transition elements
also have more

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Chapter 9 possible
oxidation state. Tin
(II) = Sn^{2+} Lead
(II) = Pb^{2+} Tin (IV)
= Sn^{4+} Lead (IV) =
 Pb^{4+}

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Assessment 42. a.

2- b. 1+ c. 1- d. 3+

43. a. 2+ b. 2+ c.

3+ d. 1+

44. a.
barium ion b.

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iodide ion e. silver
ion d. mercury(II)
ion 45. cyanide,
 CN^- and hydroxide,
 OH^- 46. a.
hydroxide ion b.
lead(IV) ion c.
sulfate ion d. oxide
ion 47. zero; A
compound is
electrically neutral.
48. The symbols for
the cation and

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Section 9.1:

Naming Ions. 1.

What is the charge
on the ion typically
formed by each
element? a. oxygen
c. sodium e. nickel,
two electrons lost .

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b. iodine d.
aluminum f.
magnesium. 2.
How many
electrons does the
neutral atom gain
or lose when each
ion forms? a. Cr^{3+}
c. Li^+ e. Cl^- b. P^{3-}
...

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In samples of any chemical compound, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ...

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