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precipitate silver iodide. chemical equation:

$KI(aq) + AgNO_3(aq) \rightarrow KNO_3(aq) + AgI(s)$  complete

ionic equation:  $K^+(aq) + I^-(aq) \rightarrow Ag^+(aq) + NO_3^-(aq)$

$K^+(aq) + NO_3^-(aq) + AgI(s)$

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(Ar = 32) weighs twice as much as an oxygen  
atom (Ar = 16), 32 g of sulfur atoms and 16 g  
of oxygen atoms both contain the same number  
of atoms. Answer: 32 g 9.2 (a) 1 mole of H  
atoms = 1 g 1 mole of H<sub>2</sub> molecules = 2 × 1 =  
2 g. (b) 1 mole of N atoms = 14 g

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2.1 namic range—are important. When a species enters a mass spectrom - eter it is ionized (the PTR—proton transfer reaction—in PTR-MS ... Chapter 2 Basic Tools of Analytical Chemistry 9 Chapter 2 1. (a) 3 significant figures; (b) 3 significant figures; (c) 5 significant fig-

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chemical names for the following: Solution a)  
Since there are four similar monodentate  
ligands. So tetra will be used as prefix. The  
metal is Cu and it's oxidation Read more...

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chemical names for the following: Solution a)  
Since there are four similar monodentate  
ligands. So tetra will be used as prefix.

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