

## Chemistry 12 Nelson Solutions

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~~Chemistry 12 - Chapter 3 Quiz. True/False. Indicate whether the sentence or statement is true or false. 1. The region in space where an electron is most~~

likely to be found is called an energy level. 2. Electron configurations are often condensed by writing them using the previous noble gas core as a starting point. ...

~~Chemistry 12 Chapter 3 Quiz Nelson~~

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Solution: OH OH OH 0.50 mmol mL -- -- $\times$  = nc V  $\times$  25.0 mL OH n - = 12.5 mmol HCl(aq) HCl(aq) [HCl(aq)] 12.5 mmol = = n V 0.10 mmol HCl(aq) /mL V = 120 mL 2. (a) Given: V HCl(aq) = 25.00 mL; [HCl(aq)] = 0.350 mol/L; V NaOH(aq) = 5.0 mL; [NaOH(aq)] = 0.500 mol/L; Required: [H<sup>+</sup>(aq)], pH Analysis: Before titration, amount of H<sup>+</sup> is: H(aq) H(aq) H(aq) 0.350 mmol mL ++ ++ $\times$  = nc V

~~Section 8.7: Acid-Base Titration Tutorial 1 Practice, page 547~~

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8 MHR Chemistry 12 Solutions Manual 978-0-07-106042-4 8. Identify any errors in the structure by drawing them. Rename the structure correctly. 2-ethyl-2,4,4-trimethylpentane What Is Required? You are to identify the errors in the name by drawing the structure and renaming it correctly. What Is Given? You are given an incorrectly named alkane.

~~Unit 1 Organic Chemistry Mr. Arthur's Science Page~~

= 0.12 mol/L Statement: The equilibrium concentration of dinitrogen tetroxide gas is 0.12 mol/L. 2. Given: initial quantity of NOCl(g) = 3.00 mol; volume = 3.00 L; [NO(g)] initial = 0.00 mol/L; [Cl<sub>2</sub>(g)] initial = 0.00 mol/L; [NO(g)] equilibrium = 0.043 mol/L Required: [NOCl(g)] equilibrium; [Cl<sub>2</sub>(g)] equilibrium Solution: Step 1.

~~Section 7.1: Equilibrium Systems Mini Investigation: The ...~~

Solution: bonds broken 3 mol 1 mol CH CCl 3 mol HnD DD --  $\Delta = \sum \text{bonds broken} - \sum \text{bonds formed} = 413 \text{ kJ mol}^{-1} \times 3 + 339 \text{ kJ mol}^{-1} \times 3 - 1578 \text{ kJ}$  Statement: The energy required to separate 1 mol of chloromethane into free atoms is 1578 kJ. 2. Solution: Step 1: Write the balanced chemical equation for the combustion of 1 mol of ethene gas.  $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$  ...

~~Section 5.3: Bond Energies Tutorial 1 Practice, page 312~~

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Solution:  $q = Ne = (1.2 \times 10^{-8}) (1.602 \times 10^{-19} \text{ C}) = 1.92 \times 10^{-11} \text{ C}$  (one extra digit carried)  $F E = kq_1 q_2 / r^2 = 8.99 \times 10^9 \text{ N m}^2 \text{ C}^{-2} \# \$ \% \& ' ((1.92 \times 10^{-11} \text{ C})^2 / (1.10 \text{ m})^2 = 2.7 \times 10^{-12} \text{ N}$  Statement: The force of repulsion between the two plastic spheres is  $2.7 \times 10^{-12} \text{ N}$ . 2. Given:  $m = 2.48 \times 10^{-15} \text{ kg}$ ;  $\Delta d = 1.7 \text{ cm} = 0.017 \text{ m}$ ;  $\Delta V_b = 260 \text{ V}$ ;  $e = 1.602 \times 10^{-19} \text{ C}$ ;  $g = 9.8 \text{ m/s}^2$  Required:  $q$ ;  $N$

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