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Q. The Quantum Mechanical Model of the Atom describes the electron's probable location around the nucleus in a 3-D cloud called a(n) \_\_\_\_.

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Electrons are found in certain orbits located around the nucleus. Every electron has a fixed energy in certain energy levels: they are said to be quantized. Electrons farther away from the nucleus have higher energy than electrons closer to the nucleus. Energy levels are closer together the farther away from the nucleus.

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The orbital diagram for a ground-state nitrogen atom is  $1s^2 2s^2 2p^3$  A: 5. The number of orbitals in a d subshell is A: 5 6. The maximum number of electrons that can occupy an energy level described by the principal quantum number, n, is A:  $2n^2$  7. A ground-state atom of manganese has \_\_\_ unpaired electrons and is \_\_\_\_\_. A: 5, paramagnetic 8.

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electrons exhibit properties of both particles and waves. d. the chemical properties of elements can be grouped according to periodicity but physical properties cannot. \_\_\_\_ 28. Elements in a group or column in the periodic table can be expected to have similar. ... Chemistry Chapter 5 Exam ...

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Chapter 5 Bond Polarity ... In Chemistry, resonance is a way of describing delocalized electrons in an atom or molecule that cannot be represented with a single Lewis dot structure. ... · Determine the total # of valence electrons in a molecule: N 5. 30 18. Ve ...

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