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How to Write the Electron

Configuration for an Element in Each

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The Periodic Table: Atomic Radius,  
Ionization Energy, and  
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& orbitals Example- Determining  
if a set of 4 Quantum Numbers is  
allowed or not ~~Quantum Numbers~~

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~~Pattern n, l, m, \u0026s on the Periodic~~

~~Table~~ Energy Levels, Energy

Sublevels, Orbitals, \u0026amp; Pauli

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Explain giving reasons, which of the following sets of quantum numbers are not possible?

~~Quantum Numbers  
(tagalog) | NOW I KNOW |~~

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~~Quantum numbers - allowable values  
identify which sets of quantum  
numbers are valid for an electron.  
each set is ordered  $(n, l, m_l, m_s)$ .~~

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~~Quantum Numbers Answers~~  
How many electrons in an atom could have these sets of quantum numbers ~~Atomic Orbitals and Quantum Numbers~~  
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chemquest, 12, quantum, numbers,

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- l □ Secondary Quantum Number/Orbital Shape Quantum number: represents the shape of the orbital- s, p, f, d. l is a range of  $n-1$ . ml
- Magnetic quantum number:

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represents the number of orbits possible.  $M_l$  is a range of  $l$ .  $m_s$  Spin Quantum number: represents the electron and its spin. Two possibilities  $+1/2$ ,

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When  $l = 0$  (s cloud),  $1$  (p cloud),  $2$  (d cloud),  $3$  (f cloud) Third  $\square$  Magnetic Quantum number ( $m_l$ ) = location or spatial orientation  $m_l = -l$  to  $l$  Fourth  $\square$  Spin Quantum number ( $m_s$ ) = electron spin  $m_s = +1/2$  or  $-1/2$ . 2. Name the orbitals described by the following quantum numbers a.  $n = 3, l = 0$  3s. c.

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