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Ionization energy of atoms, denoted E I, is measured by finding the minimal energy of light quanta or electrons accelerated to a known energy that will kick out the least bound atomic electrons.The measurement is performed in the gas phase on single atoms. While only noble gases occur as monoatomic gases, other gases can be split into single atoms.

Ionization energy - Wikipedia

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a. Which of the electron transitions involves the most energy? n=6to n=2, violet light, 410 nmz. b. Explain why this transition involves the most energy based on your understanding of the attractive forces between the electrons and protons in the atom. As the electron moves toward the nucleus of the atom there is a stronger attraction. This releases

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Cord Extension Socket – Electron Energy

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