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Answers

Empirical and

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Worksheet ANSWER

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KEY. Write the empirical formula for the following compounds.

1) C_6H_6
 CH_4 6) C_8H_{18} C_4H_9
7) WO_2 WO_2 8)
 $C_2H_6O_2$ CH_3O 9)
 $X_{39}Y_{13}$ X_3Y_6 6) A compound with an empirical formula of $C_2O_2H_4$ and a molar mass of 88 grams per mole. What is the molecular formula of

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this compound? C. 4. O.
2. H. 8. 7) A compound
with an empirical
formula of C. 4. H. 4. O
and a molar mass of 136
grams per mole.

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Students will learn simple techniques on how to solve problems involving empirical formula and molecular formula from this worksheet.

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23, 2020. Some of the worksheets below are Empirical Formula Worksheets, several exam style questions like determine the empirical formula of a compound that contains 53.70% iron and 46.30% sulfur, ... solutions are provided at the end of each worksheet.

Empirical Formula

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Worksheet #8 Empirical
Formulas 1. State the
empirical formula for
each of the following

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compounds: a) C_4H_8 ;
b) $C_2H_6O_2$; c) N_2O_5 ;
d) $Ba_3(PO_4)_2$; e) TeI_6
2. What is the

empirical formula for a
compound that contains
0.063 mol chlorine and
0.22 mol

Worksheet #8 Empirical
Formulas H O N O 4I
EMPIRICAL AND
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FORMULA

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WORKSHEET An oxide of chromium is found to have the following %

composition: 68.4 % Cr and 31.6 % O.

Determine this compound's empirical formula. The percent composition of a compound was found to be 63.5 % silver, 8.2 % nitrogen, and 28.3 % oxygen. Determine the

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compound's empirical
formula.

Worksheet

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Molecular Formula.

Empirical Formula.

Formaldehyde. CH_2

O . CH_2O . Acetic

Acid. $\text{C}_2\text{H}_4\text{O}_2$. CH

2O . Glucose. C_6H_{12}

O_6 . CH_2O

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The empirical formula
of a compound is CH_2 ,
Its molecular mass is 70
g/mol. What is its
molecular formula? =

XCI) 70 C H A
compound is found to
be 40.00% carbon,

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6.7% hydrogen and
53.5% oxygen. Its
molecular mass is 60.
g/mol, What is its
molecular formula? 401)
c 3133 4 A compound is
64.9% carbon, 13.5%
hydrogen and 21.60/0
oxygen. Its molecular

DETERMINING

EMPIRICAL

FORMULAS Name

What is the empirical ...

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The empirical formula of benzene is CH (its molecular formula is C₆ H₆). If 10.00 mg of benzene is subjected to combustion analysis, what mass of CO₂ and H₂O will be produced? Answer: The empirical formula is C₄ H₅. (The molecular formula of xylene is actually C₈ H₁₀.)
33.81 mg of CO₂; 6.92

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Chapter 11.2: Empirical
and Molecular

Formulas - Chemistry ...

composition is as
follows, what is the
empirical and molecular
formula of serine? C =
34.95 % H= 6.844 % O
= 46.56 % N= 13.59 %
Percent Composition
and Molecular Formula

Worksheet Key

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Percent Composition
and Molecular Formula
Worksheet

Molecular formula is 2
times the empirical
formula $2 (\text{C}_2\text{OH}_4) =$
 $\text{C}_4\text{O}_2\text{H}_8$ 7) A
compound with an
empirical formula of
 $\text{C}_4\text{H}_4\text{O}$ and a molar
mass of 136 grams per
mole. What is the
molecular formula of

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this compound?

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Empirical and
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KEY. Write the
empirical formula for
the following
compounds. 1) C_6H_6
 CH . 6) C_8H_{18} C_4H_9
7) WO_2 WO_2 8)

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$C_2H_6O_2$ CH_3O_9)
 $X_{39}Y_{13}$ X_3Y_6) A
compound with an
empirical formula of C.
2. OH. 4. and a molar
mass of 88 grams per
mole. Page 1 / 3

Determining Molecular
Formulas Answer Key
The empirical formula
of a compound
represents the simplest
whole-number ratio

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between the elements that make up the compound. This 10-question practice test deals with finding empirical formulas of chemical compounds. A periodic table will be required to complete this practice test. Answers for the test appear after the final question:

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