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326 - B1+B2 Bomb Calorimetry

CalorimetryCalorimetry Examples: How
to Find Heat and Specific Heat Capacity

FSc Chemistry Book1, CH 7, LEC 10:

Glass Calorimeter Calorimetry and

enthalpy introduction | Thermodynamics |

Chemistry | Khan Academy Coffee Cup

Calorimetry S6E3—Calorimetry and

~~Specific Heat Capacity 3-L1 Specific heat~~

~~calorimetry lab~~ Determination of Water

Equivalent of a Calorimeter **SPECIFIC**

HEAT CAPACITY EXPERIMENT

Rate of Reaction of Sodium Thiosulfate

and Hydrochloric Acid Calorimetry

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RESOURCES The CCLI Initiative
Learning Objectives □ understand the
concept of heat and joules. □ perform heat-
gain and heat-loss calculations. Procedure
Overview □ hot and cold water are mixed.
The water masses and temperature
changes are used to calculate the heat lost

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INTRODUCTION TO CALORIMETRY

Report Sheet (page 3) Questions/Problems

1. Distinguish between heat and temperature.
2. Describe what is meant by the specific heat of a substance.
3. In this experiment, you measured the heat associated with a burning candle. Does this represent a physical change and/or a chemical change? Describe the change(s) occurring.

INTRODUCTION TO CALORIMETRY

Introduction to

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oscillations and waves 40

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Calorimeters are used to measure potential energy. A calorie is the energy it takes to heat 1 mL of water 1 degree Celsius.

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These calories are not the same as those used to refer to food on nutrition labels, dietary plans, etc., which are known as Calories or Kcal (1000 normal calories).

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