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24 Kinetic Molecular Theory \u0026amp; Diffusion

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### Kinetic Molecular Theory

The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 ©POGIL - 2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University

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Read Online Kinetic Molecular Theory Pogil 2005 Answers molecular theory to predict the outcome of everyday situations. POGIL: Kinetic Molecular Theory The kinetic molecular theory (KMT) is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter. This theory is based on the Page 7/25

### ~~Kinetic Molecular Theory Pogil 2005 Answers~~

The kinetic molecular theory (KMT) is a simple microscopic model that effectively explains the gas laws described in previous modules of this chapter. This theory is based on the following five postulates described here. (Note: The term “ molecule ” will be used to refer to the individual chemical species that compose the gas, although some ...

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The particles are assumed to not attract nor repel each other. The average kinetic energy of the gas particles is directly proportional to the Kelvin temperature of the gas. 3/3 ©POGIL -2005 Authored by Edited by Linda Padwa and David Hanson, Stony Brook University. Kinetic Molecular Theory Key Questions 1.

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The Kinetic-Molecular Theory {Boyle ' s Law  $zP \propto 1/V$  zAs the V increases the molecular collisions with container walls decrease and the P decreases.  
{Dalton ' s Law  $zP_{total} = P_A + P_B + P_C + \dots$  zBecause gases have few intermolecular attractions, their pressures are independent of other gases in the container. {Charles ' Law  $zV \propto T$

## ~~CHAPTER 12 GASES AND KINETIC MOLECULAR THEORY~~

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Information (Kinetic Molecular Theory) The observed behavior of gases, as expressed by the empirical gas laws, can be understood on the basis of the kinetic molecular theory, which developed over many years up to 1857, when Rudolf Clausius (1822-1888) published it in its most complete and elegant form.

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POGIL: Kinetic Molecular Theory - Studyres The Kinetic-Molecular Theory Explains the Behavior of Gases, Part II According to Graham ' s law, the molecules of a gas are in rapid motion and the molecules themselves are small. The average distance between the molecules of a gas is large compared to the size of the molecules. Page 1/5

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Kinetic Molecular Theory states that gas particles are in constant motion and exhibit perfectly elastic collisions. Kinetic Molecular Theory can be used to explain both Charles ' and Boyle ' s Laws. The average kinetic energy of a collection of gas particles is directly proportional to absolute temperature only.

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Kinetic Molecular Theory is guided by the following assumptions: The molecules of an ideal gas (an ideal gas is a gas that follows the assumptions of kinetic molecular theory) are in constant, random, straight-line motion. Gas molecules are constantly moving very, very fast and their direction is completely random.

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kinetic molecular theory answer key, Feb 19, 2003 - 2. State the six postulates of the Kinetic-molecular theory and explain how they account for the physical properties of gases. Key Terms: diffusion elastic Kinetic-molecular theory. Notes: (13-1) Nature of Gases: We are constantly surrounded and buried in gas molecules.

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