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4.2 For some hypothetical metal, the equilibrium number of vacancies at 900 C is  $2.3 \times 10^{25} \text{ m}^{-3}$ . If the density and atomic weight of this metal are  $7.40 \text{ g/cm}^3$  and  $85.5 \text{ g/mol}$ , respectively, calculate the fraction of vacancies for this metal at 900 C. Solution This problem is solved using two steps: (1) calculate the total number of lattice sites in silver,  $N$ , using

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2.2 Silicon has three naturally occurring isotopes: 92.23% of  $^{28}\text{Si}$ , with an atomic weight of  $27.9769 \text{ amu}$ , 4.68% of  $^{29}\text{Si}$ , with an atomic weight of  $28.9765 \text{ amu}$ , and 3.09% of  $^{30}\text{Si}$ , with an atomic ...

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