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grams and mL)

Conversion of
Molarity (M) to
Molality (m) by

Abhishek Jain

(ABCH Sir) for IIT
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~~| normality |~~

~~formality~~ Calculate

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percent by mass

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worksheet. Find the
molarity of the

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following solutions.

214 2g K_2SO_4 x 1 mol

K_2SO_4 129 m K_2SO_4 0

0673 l soln 247 23

g K_2SO_4 calculate the

molarity if a flask

contains 154 moles

potassium sulfate in

125 ml of solution.

Molarity by dilution

acids are usually

acquired from

chemical supply

houses in

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Worksheet

~~Molarity M~~

~~Worksheet~~

~~Answers~~

~~worksheet~~

MOLARITY (M) =
m oles of solute

MOLALITY (m or
) = m oles of

solute Liters of
solvent kg of

solvent Molarity

Example: 4.0 moles

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of LiCl is dissolved
in 5.0 liters of
water. What is the
molarity of the
solution? 4.0 moles
= 0.8 M This
solution is 0.8
Molar or 0.8 M 5.0
Liters

~~7) How many moles
of solute are in 125
mL of a 2.0 M ...
Calculate molarity~~

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of 35.0 mL KOH solution needed to completely neutralize 22.5 mL of 1.75 M H₂SO₄. Calculate volume (mL) of 2.50M H₂SO₄ needed to completely neutralize 10.0g NaOH (s).

Answers. $M_1 V_1$

$= M_2 V_2$ (1.71

M)(25.0 mL) = M_2

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$$\begin{aligned} (65.0 \text{ mL}) M_2 &= \\ 0.658 \text{ M}; M &= \text{mol/L} \\ &= (25.0/40.0) / \\ (0.325) &= 1.92 \\ &\text{mol/L} \end{aligned}$$

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What is the
molarity? 214.2g
 $\text{OsF}_3 \times 1 \text{ mol OsF}_3$
 $= 12.9 \text{ M OsF}_3$
 0.0673 L soln
 247.23 g OsF_3

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Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.
1.54...

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Worksheet #2:
Molarity and
Dilution Problems

1) Describe how
you would prepare
5.00 liters of a
6.00M solution of
potassium
hydroxide.

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Solutions

~~Worksheet 2~~

~~Molarity And~~

~~Dilution Problems ...~~

Molarity and

Molality Practice

Worksheet. Find the

molarity of the

following solutions:

- 1) 0.5 moles of sodium chloride is dissolved to make 0.05 liters of

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solution. 2) 0.5
grams of sodium
chloride is
dissolved to make
0.05 liters of ...
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problems

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Answers: 1.

39.2-ml (Put in paragraph form) 2.

1.1 M 3. 4.64 M 4.

49.2-ml Take

49.2-ml of 18.0 M

H₂SO₄ stock

solution and pour it into a 500-ml

volumetric flask.

Fill to the 500-ml

line with distilled

water to make

1.77M H₂SO₄

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solution. Extra
Molarity Problems
for Practice 1. How
many moles of LiF
would be required
to produce a 2.5 M
solution with a

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molarity practice
worksheet answers,
I note that when we

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do our lab we will be measuring in ml, so this is good practice. The final problem should be set up as $0.192 \text{ moles} / 0.450 \text{ L}$, which gives us $M = 0.43$. I then instruct students to complete problems 4-11 on the Molarity and Neutralization

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This quiz and worksheet allow students to test the following skills:

Problem solving - use acquired knowledge to answer practice problems involving the calculation of molality

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Molality Worksheet

Complete the

following questions
and problems
relating to molality.

- 1) Write the equation for molality:
- 2) Write the equation for molarity:
- 3) Explain in words how molality and molarity differ.
- 4)

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What is the molality of a solution in which 0.32 moles aluminum chloride has been dissolved in 2,200 g water?

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Molarity = Moles of
solute / Liters of
Solution

(abbreviation = M)

Molality = Moles of
solute / Kg of
Solvent

(abbreviation = m)

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Normality =
number of
equivalent of solute
x Molarity of
Solution

(abbreviation = N)

Mass Percent =
mass of solute /
mass of solution. 1.

How many moles of
ethyl alcohol, C.

~~Honors Chemistry~~

~~Name Chapter 12:~~

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-
Answers
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o 1) Molality is
moles solute
dissolved per
kilogram of solvent.

2) Let moles of
solute be
represented by 'n'.

3) The formula for
acetone is C_3H_6O
and its molar

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mass is 58.0794
g/mol, which equals
0.0580794 kg/mol.

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