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Mole Ratios from Balanced Equations
Stoichiometry Mole to Mole Ratio
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Limiting Reactant Practice Problem

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Practice

Calculating Moles in a Balanced Equation
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use balanced chemical equations? Why?

Chemists use balanced chemical equations
as a basis to calculate how much reactant

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is needed or product is formed in a reaction. This is called Stoichiometry- (stoi-key-ah-meh-tree) Another way of looking at it is using the mole ratio from the balanced equation and information about ...

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5.2 g O₂ * 1 mol O₂ * 2 mol H₂ * 22.4 L

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$H_2 = 7.28 \text{ L } H_2$. $32 \text{ g } O_2$ $1 \text{ mol } O_2$ 1 mol

H_2 . 17. Discuss as a group the process to convert between different chemical species in a given reaction. Make sure you address how the new mole map can be used in this process. Using complete.

HS Chemistry POGIL Activity

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Mole Ratio Worksheet. 1) Balance this equation: $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$, write the following molar ratios: N_2 / H_2 N_2 / NH_3 H_2 / NH_3 . 2) Balance this equation: $\text{H}_2 + \text{S} \rightarrow \text{H}_2\text{S}$, write the following molar ratios: $\text{H}_2 / \text{H}_2\text{S}$ H_2 / S $\text{H}_2\text{S} / \text{S}$. 3) Write and balance the equation for the synthesis of water.

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Just as one mole of atoms equals the atomic mass of an atom in grams, so also one mole of molecules equals the molar mass of the molecule in grams. Molar mass is the mass (in grams per mole) of

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one mole of a substance.

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Activities for ... Use mathematical concepts to explain how your answer in part a is possible. 6. The ratio obtained from the coefficients in a balanced

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Chemical equation is called the mole ratio.

- a. What is the mole ratio for the reaction in Model 1? b.

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Mole ratio Answer: 34.8 3 mol H =
52.2 mol H₂ 2 mol NH₃ Key Questions: 1.
What was the given in the example above?
3 mol H₂ 2. Where did the come from? 2
mol NH₃' OR Mole ratio 34.8 mol x 3 mol
= H₂ 2 mol NH₃ 3. Double check to see if
52.2 moles of H₂ is the correct answer.
SHOW the math below.

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